AMENDMENTS TO THE SPECIFICATION

Please amend the paragraph beginning on line 11 of page 1 as follows:

Polyesters produced by the polymerization of 2,6-naphthalene dicarboxylic acid ~~and~~ with a diol ~~have been~~ are known to ~~be excellent~~ excel in several properties such as thermal stability, tensile strength, gas permeability, etc., and ~~they are expected to be used~~ as such are chosen as good materials for films, fibers, storage containers, etc. In particular, polyethylene naphthalate (PEN) produced by the polymerization of 2,6-naphthalene dicarboxylic acid and ethylene glycol is expected to replace polyethylene terephthalate (PET).

Please amend the paragraph beginning on line 18 of page 1 as follows:

2,6-Naphthalene dicarboxylic acid is obtained by oxidizing 2,6-dimethyl naphthalene with oxygen gas in the presence of ~~the catalysts of~~ catalytic cobalt, manganese, and bromine compounds. The thus-obtained crude 2,6-naphthalene dicarboxylic acid contains numerous impurities including acids having one functional group such as formyl naphthoic acid, methyl naphthoic acid, etc. that are generated by incomplete oxidation of 2,6-dimethyl naphthalene; trimellitic acid that is obtained from the collapse of naphthalene structure; and brominated naphthalene dicarboxylic acid, naphthoic acid, naphthalene tricarboxylic acid, colored organic impurities whose structures are not identified, and metal impurities such as cobalt complex, manganese complex, etc.

Please amend the paragraph beginning on line 8 of page 2 as follows:

Polyesters obtained by the polymerization of ethylene glycol and crude 2,6-naphthalene dicarboxylic acid containing ~~many~~ impurities as mentioned above show poor physical properties,

heat stability, structural stability, and so on. Moreover, such polyesters are classified as low quality because they are tinged with color.

Please amend the paragraph beginning on line 13 of page 2 as follows:

Of the impurities, mono carboxylic acids such as methyl naphthoic acid, naphthoic acid, etc. are especially problematic. If these mono carboxylic acids exceed a certain amount, the polymerization rate is decreased during the production of polyesters, and gelation and coloring occur. In particular, formyl naphthoic acid ~~has a serious effect on such problems~~ is especially problematic. Accordingly, so as to obtain polyesters having high quality, it is important to reduce these impurities.

Please amend the paragraph beginning on line 17 of page 3 as follows:

As a known method for refining 2,6-naphthalene dicarboxylic acid by dissolving it in general solvents and then recrystallizing it, US Patent No. 5,256,817 discloses a method for refining 2,6-naphthalene dicarboxylic acid by dissolving it in water or acetic acid and then hydrogenating and crystallizing it. However, as this method requires heating to a high temperature in order to dissolve the 2,6-naphthalene dicarboxylic acid, the production cost of naphthoic acid is increased and it also requires expensive metal catalysts for hydrogenation, and is thus problematic.

Please amend the paragraph beginning on line 3 of page 15 as follows:

To a 1-neck Erlenmeyer flask having a Pyrex-type lid, 30.0 g of crude 2,6-naphthalene dicarboxylic acid and 33.4 g of triethyl amine were added at room temperature and room

pressure. At 50°C, 60 g of methanol were added to the above mixture, which was then stirred for 30 minutes to obtain a solution of the amine salt of 2,6-naphthalene dicarboxylic acid. After the amine salt solution was filtrated using a filter with a 7-μm pore size, 240 g of methyl acetate was added to the filtrate to mix them, and the resultant was then cooled at 0°C for 12 hours. The amine salt crystal of 2,6-naphthalene dicarboxylic acid obtained after cooling was filtrated and ~~deaminized~~ deaminated at 90°C to yield a purified 2,6-naphthalene dicarboxylic acid.

Please amend the paragraph beginning on line 3 of page 16 as follows:

After 30.0 g of crude 2,6-naphthalene dicarboxylic acid and 33.4 g of triethyl amine were added to a 4-neck Erlenmeyer flask having a Pyrex-type lid at room temperature and room pressure, 315 g of a mixed solution containing methanol and methyl acetate in a mixing ratio of 2:8 were added thereto and the mixture was heated to 55°C while stirring for 30 minutes to obtain a solution of the amine salt of 2,6-naphthalene dicarboxylic acid. The amine salt solution was placed at room temperature for 12 hours to thereby crystallize it, and the thus-produced diamine salt crystal of 2,6-naphthalene dicarboxylic acid was then filtrated and ~~deaminized~~ deaminated at 90°C to yield a refined 2,6-naphthalene dicarboxylic acid powder.

Please amend the paragraph beginning on line 2 of page 20 as follows:

After 50.0 g of crude 2,6-naphthalene dicarboxylic acid and 56.5 g of triethyl amine were added to a 4-neck Erlenmeyer flask having a Pyrex-type lid at room temperature and room pressure, 315 g of a mixed solution containing methanol:water:methyl acetate in a mixing ratio of 1.75:0.25:8.00 were added thereto and the mixture was then heated to 55°C while stirring for 1 hour to obtain a solution of the amine salt of 2,6-naphthalene dicarboxylic acid. The amine salt

solution was filtrated using a filter with a 7 μm pore size at 60°C under a reduced pressure, the thus-obtained filtrate was heated to 55°C for 30 minutes to convert it into its solution form, and then it was placed for 12 hours to thereby crystallize it. The thus-produced diamine salt crystal of 2,6-naphthalene dicarboxylic acid was filtrated and ~~deaminized~~ deaminated at 90°C to yield a refined 2,6-naphthalene dicarboxylic acid powder.